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Solutions 1) If I initially have 4.0 L of a gas at a pressure of 1.1 atm, what will the volume be if I increase the pressure to 3.4 atm? $(1.1 \text{ atm})(4.0 \text{ L}) = (3.4 \text{ atm})(x \text{ L})$ $x = 1.29 \text{ L}$ 2) A toy balloon has an internal pressure of 1.05 atm and a volume of 5.0 L.

Combined Gas Law Worksheet

Combined Gas Law Chart Worksheet

Answers Combined gas law formula:

$PV/T = k$. Where: $k = \text{constant}$. $P = \text{pressure}$. $V = \text{Volume}$. $T = \text{temperature}$.

In order to compute the changes in temperature, pressure or volume a sample gas may suffer in certain

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conditions, the combined gas law can be written in the form detailed within the next rows: $P_1 V_1 / T_1 \dots$

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Combined Gas Law Chart Combined Gas Law Worksheet - Solutions 1) If I initially have 4.0 L of a gas at a pressure of 1.1 atm, what will the volume be if I increase the pressure to 3.4 atm? $(1.1 \text{ atm})(4.0 \text{ L}) = (3.4 \text{ atm})(x \text{ L})$ $x = 1.29 \text{ L}$ 2) A toy balloon has an internal pressure of 1.05 atm and a volume of 5.0 L. Combined Gas Law Worksheet

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icdovidiocb.gov.it our books like this one. [MOBI] Combined Gas Law Chart Worksheet Answers Combined Gas Law Chart Answer Key Wed, 22 Jul 2020 14:47 Combined Gas Law Problems: 1 atm = 760.0 mm Hg = 101.3 kPa k = 273 + °C A gas balloon has a volume of 106.0 liters when the temperature is 45.0 °C

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Combined Gas Law Chart Worksheet Answers P_1 = Initial Pressure ; V_1 =

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Initial Volume ; T_1 = Initial Temperature ; P_2 = Final Pressure ; V_2 = Final Volume ; T_2 = Final Temperature.

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increase the pressure to 3.4 atm?

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Atmospheric Pressure: 1 atm = 760 torr
= 760 mm Hg = 101.3 kPa = 14.7 psi
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